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## Structure Reports

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## 2,2'-Hexamethylenedi-1,3-benzothiazole

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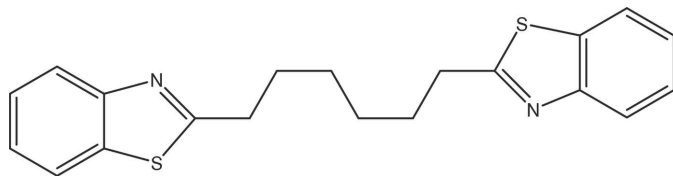
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Key indicators: single-crystal X-ray study;  $T = 293$  K; mean  $\sigma(\text{C}-\text{C}) = 0.006$  Å;  $R$  factor = 0.062;  $wR$  factor = 0.182; data-to-parameter ratio = 14.9.

The title compound,  $\text{C}_{20}\text{H}_{20}\text{N}_2\text{S}_2$ , was prepared by the reaction of suberic acid and 2-aminothiophenol under microwave irradiation. The molecule lies on an inversion center.

## Related literature

For details of the synthesis and the application of benzothiazoles, see: Chakraborti *et al.* (2004); Seijas *et al.* (2007); Wang *et al.* (2009). For the use of microwave-assisted organic synthesis, see: Kappe & Stadler (2005).



## Experimental

## Crystal data

$\text{C}_{20}\text{H}_{20}\text{N}_2\text{S}_2$   
 $M_r = 342.50$

Monoclinic,  $P2_1/n$   
 $a = 5.7590$  (12) Å

$b = 8.3030$  (17) Å  
 $c = 18.974$  (4) Å  
 $\beta = 96.03$  (3)°  
 $V = 902.3$  (3) Å<sup>3</sup>  
 $Z = 2$

Mo  $K\alpha$  radiation  
 $\mu = 0.30$  mm<sup>-1</sup>  
 $T = 293$  (2) K  
 $0.30 \times 0.20 \times 0.10$  mm

## Data collection

Enraf–Nonius CAD-4 diffractometer  
 Absorption correction:  $\psi$  scan (North *et al.*, 1968)  
 $T_{\min} = 0.916$ ,  $T_{\max} = 0.971$   
 1626 measured reflections

1626 independent reflections  
 1102 reflections with  $I > 2\sigma(I)$   
 3 standard reflections every 200 reflections  
 intensity decay: 9%

## Refinement

$R[F^2 > 2\sigma(F^2)] = 0.062$   
 $wR(F^2) = 0.182$   
 $S = 1.01$   
 1626 reflections

109 parameters  
 H-atom parameters constrained  
 $\Delta\rho_{\max} = 0.37$  e Å<sup>-3</sup>  
 $\Delta\rho_{\min} = -0.40$  e Å<sup>-3</sup>

Data collection: *CAD-4 Software* (Enraf–Nonius, 1989); cell refinement: *CAD-4 Software*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 2008); program(s) used to refine structure: *SHELXL97* (Sheldrick, 2008); molecular graphics: *SHELXTL* (Sheldrick, 2008); software used to prepare material for publication: *SHELXTL*.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: BX2191).

## References

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**supplementary materials**

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## 2,2'-Hexamethylenedi-1,3-benzothiazole

G. Wang, L. Zhuang and J. Wang

### Comment

Benzothiazole are remarkable heterocyclic ring systems. They have been found to exhibit a wide spectrum of biological activities. Many kinds of 2-substituted benzothiazoles are utilized as vulcanization accelerators in the manufacture of rubber, as fluorescent brightening agents in textile dyeing, and in the leather industry (Chakraborti *et al.*, 2004; Seijas *et al.*, 2007; Wang *et al.*, 2009). There are numerous synthetic methods to produce 2-arylbenzothiazoles. The most important ones include the reaction of *o*-aminothiophenols with benzoic acids or their derivatives (Chakraborti *et al.*, 2004; Seijas *et al.*, 2007; Wang *et al.*, 2009). Microwave-assisted organic synthesis (MAOS) is a powerful technique that is being used more and more to accelerate thermal organic reactions (Kappe & Stadler, 2005). We are focusing on Microwave-assisted synthesis of new products of bisbenzothiazole. We here report the crystal structure of the title compound (I). The atom-numbering scheme of (I) is shown in Fig. 1. The compound lies on an inversion center (symmetry code  $-x+1, -y, -z$ ).

### Experimental

A mixture of 2-aminothiophenol (2.5 g, 20 mmol), 5 ml orthophosphoric acid, 5 g polyphosphoric acid and 1,6-hexanedicarboxylic acid (1.74 g, 10 mmol) in a beakerflask (150 ml) was placed in a domestic microwave oven (0.8 KW, 2450 MHz) and irradiated (micromode, full power) for 4 min (30 s per time). The reaction mixture was cooled to r.t. and washed with aq NaOH (20%, 150 ml). The pH was adjusted to 10, the resulted solide was filtered. Then the crude compound (I) was obtained. It was crystallized from ethanol. Crystals of (I) suitable for X-ray diffraction were obtained by slow evaporation of methanol.  $^1\text{H NMR}$  (DMSO,  $\delta$ , p.p.m.) 7.35–7.40 (m, 2 H), 7.46–7.51 (m, 2 H), 7.64 (dd, 2 H), 7.81 (d, 2 H), 7.95 (dd, 2 H), 8.05 (d, 2 H).

### Refinement

All H atoms were positioned geometrically, with C—H = 0.93 and 0.97 Å for methyl and methylene H atoms, and constrained to ride on their parent atoms, with  $U_{\text{iso}}(\text{H}) = xU_{\text{eq}}(\text{C})$ , where  $x = 1.5$  for methyl H and  $x = 1.2$  for methylene H atoms.

### Figures

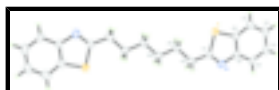


Fig. 1. A view of the molecular structure of (I) showing the atom-numbering scheme and 30% displacement ellipsoids. Unlabeled atoms are related to labeled atoms by symmetry code ( $-x+1, -y, -z$ ).

## 2,2'-Hexamethylenedi-1,3-benzothiazole

### Crystal data

$\text{C}_{20}\text{H}_{20}\text{N}_2\text{S}_2$

$F_{000} = 372$

# supplementary materials

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$M_r = 342.50$

Monoclinic,  $P2_1/n$

Hall symbol: -P 2yn

$a = 5.7590$  (12) Å

$b = 8.3030$  (17) Å

$c = 18.974$  (4) Å

$\beta = 96.03$  (3)°

$V = 902.3$  (3) Å<sup>3</sup>

$Z = 2$

$D_x = 1.297$  Mg m<sup>-3</sup>

Mo  $K\alpha$  radiation

$\lambda = 0.71073$  Å

Cell parameters from 27 reflections

$\theta = 1-25^\circ$

$\mu = 0.30$  mm<sup>-1</sup>

$T = 293$  (2) K

Block, yellow

$0.30 \times 0.20 \times 0.10$  mm

## Data collection

Enraf-Nonius CAD-4  
diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 293$ (2) K

$\omega/2\theta$  scans

Absorption correction:  $\psi$  scan  
(North *et al.*, 1968)

$T_{\min} = 0.916$ ,  $T_{\max} = 0.971$

1626 measured reflections

1626 independent reflections

1102 reflections with  $I > 2\sigma(I)$

$R_{\text{int}} = 0.0000$

$\theta_{\max} = 25.3^\circ$

$\theta_{\min} = 2.2^\circ$

$h = -6 \rightarrow 6$

$k = 0 \rightarrow 9$

$l = 0 \rightarrow 22$

3 standard reflections

every 200 reflections

intensity decay: 9%

## Refinement

Refinement on  $F^2$

Least-squares matrix: full

$R[F^2 > 2\sigma(F^2)] = 0.062$

$wR(F^2) = 0.182$

$S = 1.01$

1626 reflections

109 parameters

Primary atom site location: structure-invariant direct  
methods

Secondary atom site location: difference Fourier map

Hydrogen site location: inferred from neighbouring  
sites

H-atom parameters constrained

$$w = 1/[\sigma^2(F_o^2) + (0.06P)^2 + 1.95P]$$

where  $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} < 0.001$

$\Delta\rho_{\max} = 0.37$  e Å<sup>-3</sup>

$\Delta\rho_{\min} = -0.39$  e Å<sup>-3</sup>

Extinction correction: none

## Special details

**Geometry.** All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

**Refinement.** Refinement of  $F^2$  against ALL reflections. The weighted  $R$ -factor  $wR$  and goodness of fit  $S$  are based on  $F^2$ , conventional  $R$ -factors  $R$  are based on  $F$ , with  $F$  set to zero for negative  $F^2$ . The threshold expression of  $F^2 > \sigma(F^2)$  is used only for calculating  $R$ -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement.  $R$ -factors based on  $F^2$  are statistically about twice as large as those based on  $F$ , and  $R$ -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters ( $\text{\AA}^2$ )

|      | $x$          | $y$          | $z$          | $U_{\text{iso}}^*/U_{\text{eq}}$ |
|------|--------------|--------------|--------------|----------------------------------|
| S    | 0.21252 (19) | 0.52096 (15) | 0.08482 (6)  | 0.0614 (4)                       |
| N    | 0.6060 (6)   | 0.5674 (4)   | 0.16015 (18) | 0.0569 (9)                       |
| C1   | 0.0926 (8)   | 0.8256 (6)   | 0.1352 (3)   | 0.0701 (13)                      |
| H1A  | -0.0542      | 0.8244       | 0.1096       | 0.084*                           |
| C2   | 0.1629 (9)   | 0.9538 (6)   | 0.1788 (3)   | 0.0738 (14)                      |
| H2A  | 0.0601       | 1.0387       | 0.1831       | 0.089*                           |
| C3   | 0.3826 (9)   | 0.9590 (6)   | 0.2162 (2)   | 0.0673 (13)                      |
| H3A  | 0.4266       | 1.0481       | 0.2442       | 0.081*                           |
| C4   | 0.5351 (8)   | 0.8342 (5)   | 0.2123 (2)   | 0.0573 (11)                      |
| H4A  | 0.6813       | 0.8370       | 0.2382       | 0.069*                           |
| C5   | 0.4707 (7)   | 0.7038 (5)   | 0.1695 (2)   | 0.0468 (9)                       |
| C6   | 0.2468 (7)   | 0.6990 (5)   | 0.1308 (2)   | 0.0539 (10)                      |
| C7   | 0.4953 (7)   | 0.4647 (5)   | 0.1187 (2)   | 0.0495 (9)                       |
| C8   | 0.5905 (8)   | 0.3044 (5)   | 0.1000 (2)   | 0.0631 (12)                      |
| H8A  | 0.7463       | 0.3214       | 0.0864       | 0.076*                           |
| H8B  | 0.6065       | 0.2391       | 0.1426       | 0.076*                           |
| C9   | 0.4567 (8)   | 0.2091 (5)   | 0.0429 (2)   | 0.0556 (10)                      |
| H9A  | 0.4414       | 0.2725       | -0.0002      | 0.067*                           |
| H9B  | 0.3009       | 0.1901       | 0.0562       | 0.067*                           |
| C10  | 0.5656 (8)   | 0.0488 (5)   | 0.0278 (2)   | 0.0583 (11)                      |
| H10A | 0.7206       | 0.0683       | 0.0139       | 0.070*                           |
| H10B | 0.5836       | -0.0134      | 0.0712       | 0.070*                           |

Atomic displacement parameters ( $\text{\AA}^2$ )

|     | $U^{11}$   | $U^{22}$   | $U^{33}$   | $U^{12}$     | $U^{13}$    | $U^{23}$     |
|-----|------------|------------|------------|--------------|-------------|--------------|
| S   | 0.0522 (7) | 0.0655 (8) | 0.0654 (7) | -0.0031 (5)  | 0.0018 (5)  | -0.0103 (6)  |
| N   | 0.063 (2)  | 0.056 (2)  | 0.051 (2)  | -0.0014 (17) | 0.0036 (16) | -0.0028 (17) |
| C1  | 0.061 (3)  | 0.075 (3)  | 0.076 (3)  | 0.008 (2)    | 0.018 (2)   | 0.002 (3)    |
| C2  | 0.084 (4)  | 0.065 (3)  | 0.078 (3)  | 0.013 (3)    | 0.035 (3)   | -0.007 (3)   |
| C3  | 0.088 (4)  | 0.061 (3)  | 0.058 (3)  | -0.011 (3)   | 0.031 (2)   | -0.013 (2)   |
| C4  | 0.067 (3)  | 0.063 (3)  | 0.043 (2)  | -0.014 (2)   | 0.0089 (19) | -0.011 (2)   |
| C5  | 0.057 (2)  | 0.044 (2)  | 0.040 (2)  | -0.0024 (17) | 0.0089 (17) | 0.0032 (17)  |
| C6  | 0.050 (2)  | 0.066 (3)  | 0.047 (2)  | -0.006 (2)   | 0.0127 (18) | -0.006 (2)   |
| C7  | 0.055 (2)  | 0.047 (2)  | 0.046 (2)  | -0.0056 (18) | 0.0063 (17) | -0.0045 (18) |
| C8  | 0.070 (3)  | 0.053 (3)  | 0.066 (3)  | 0.009 (2)    | 0.005 (2)   | 0.007 (2)    |
| C9  | 0.067 (3)  | 0.052 (2)  | 0.048 (2)  | -0.002 (2)   | 0.0086 (19) | 0.0016 (19)  |
| C10 | 0.071 (3)  | 0.053 (2)  | 0.052 (2)  | 0.007 (2)    | 0.012 (2)   | 0.002 (2)    |

## supplementary materials

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### Geometric parameters (Å, °)

|             |            |                            |            |
|-------------|------------|----------------------------|------------|
| S—C6        | 1.717 (4)  | C4—H4A                     | 0.9300     |
| S—C7        | 1.750 (4)  | C5—C6                      | 1.416 (5)  |
| N—C7        | 1.283 (5)  | C7—C8                      | 1.496 (6)  |
| N—C5        | 1.397 (5)  | C8—C9                      | 1.489 (6)  |
| C1—C2       | 1.382 (7)  | C8—H8A                     | 0.9700     |
| C1—C6       | 1.384 (6)  | C8—H8B                     | 0.9700     |
| C1—H1A      | 0.9300     | C9—C10                     | 1.512 (6)  |
| C2—C3       | 1.385 (7)  | C9—H9A                     | 0.9700     |
| C2—H2A      | 0.9300     | C9—H9B                     | 0.9700     |
| C3—C4       | 1.365 (6)  | C10—C10 <sup>i</sup>       | 1.473 (8)  |
| C3—H3A      | 0.9300     | C10—H10A                   | 0.9700     |
| C4—C5       | 1.380 (5)  | C10—H10B                   | 0.9700     |
| C6—S—C7     | 89.46 (19) | N—C7—S                     | 115.5 (3)  |
| C7—N—C5     | 111.6 (4)  | C8—C7—S                    | 120.0 (3)  |
| C2—C1—C6    | 118.1 (5)  | C9—C8—C7                   | 118.1 (4)  |
| C2—C1—H1A   | 120.9      | C9—C8—H8A                  | 107.8      |
| C6—C1—H1A   | 120.9      | C7—C8—H8A                  | 107.8      |
| C1—C2—C3    | 121.7 (5)  | C9—C8—H8B                  | 107.8      |
| C1—C2—H2A   | 119.2      | C7—C8—H8B                  | 107.8      |
| C3—C2—H2A   | 119.2      | H8A—C8—H8B                 | 107.1      |
| C4—C3—C2    | 120.4 (4)  | C8—C9—C10                  | 114.4 (4)  |
| C4—C3—H3A   | 119.8      | C8—C9—H9A                  | 108.7      |
| C2—C3—H3A   | 119.8      | C10—C9—H9A                 | 108.7      |
| C3—C4—C5    | 119.5 (4)  | C8—C9—H9B                  | 108.7      |
| C3—C4—H4A   | 120.2      | C10—C9—H9B                 | 108.7      |
| C5—C4—H4A   | 120.2      | H9A—C9—H9B                 | 107.6      |
| C4—C5—N     | 126.3 (4)  | C10 <sup>i</sup> —C10—C9   | 115.4 (5)  |
| C4—C5—C6    | 120.1 (4)  | C10 <sup>i</sup> —C10—H10A | 108.4      |
| N—C5—C6     | 113.6 (4)  | C9—C10—H10A                | 108.4      |
| C1—C6—C5    | 120.1 (4)  | C10 <sup>i</sup> —C10—H10B | 108.4      |
| C1—C6—S     | 130.1 (4)  | C9—C10—H10B                | 108.4      |
| C5—C6—S     | 109.7 (3)  | H10A—C10—H10B              | 107.5      |
| N—C7—C8     | 124.4 (4)  |                            |            |
| C6—C1—C2—C3 | 1.3 (7)    | N—C5—C6—S                  | 0.9 (4)    |
| C1—C2—C3—C4 | -1.6 (7)   | C7—S—C6—C1                 | -179.0 (5) |
| C2—C3—C4—C5 | 1.3 (7)    | C7—S—C6—C5                 | -0.5 (3)   |
| C3—C4—C5—N  | -179.7 (4) | C5—N—C7—C8                 | -178.0 (4) |
| C3—C4—C5—C6 | -0.8 (6)   | C5—N—C7—S                  | 0.5 (5)    |
| C7—N—C5—C4  | 178.0 (4)  | C6—S—C7—N                  | 0.0 (3)    |
| C7—N—C5—C6  | -0.9 (5)   | C6—S—C7—C8                 | 178.6 (4)  |
| C2—C1—C6—C5 | -0.8 (7)   | N—C7—C8—C9                 | -170.9 (4) |
| C2—C1—C6—S  | 177.5 (4)  | S—C7—C8—C9                 | 10.6 (6)   |
| C4—C5—C6—C1 | 0.5 (6)    | C7—C8—C9—C10               | -179.9 (4) |
| N—C5—C6—C1  | 179.6 (4)  | C8—C9—C10—C10 <sup>i</sup> | 179.1 (5)  |
| C4—C5—C6—S  | -178.1 (3) |                            |            |

Symmetry codes: (i)  $-x+1, -y, -z$ .

Fig. 1

